

Aグループ(2サンプル)

るよ、電位は増加(電位の絶対値は減少)して、



Na content (ppm)

Figure 1 shows the plot of $\log \frac{1}{1 - \alpha}$ versus $\log t$ for the polymerization of styrene in benzene solution at 50°C. The data points are shown in Figure 1, and the straight line indicates that the reaction follows first-order kinetics. The slope of the line is 0.2, which is the order of reaction with respect to styrene. The intercept of the line on the y-axis is 0.1, which is the order of reaction with respect to the initiator.

The rate of polymerization is given by the equation $R_p = k_p [M]^{0.2} [I]^{0.1}$. The rate constant k_p is determined from the slope of the line in Figure 1. The rate constant k_p is $1.5 \times 10^{-4} \text{ min}^{-1}$. The rate constant k_p is $1.5 \times 10^{-4} \text{ min}^{-1}$.

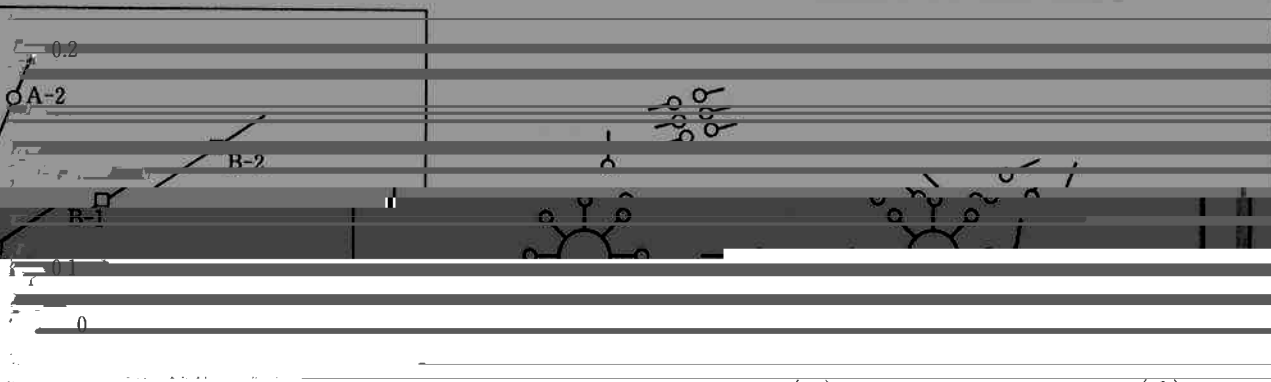


Fig. 1. The plot of $\log \frac{1}{1 - \alpha}$ versus $\log t$ for the polymerization of styrene in benzene solution at 50°C.

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[REDACTED]