

ΔH_f° (kJ/mol) for $\text{NaHCO}_3(s)$ is -933.0
 ΔH_f° (kJ/mol) for $\text{Na}(s)$ is 0
 ΔH_f° (kJ/mol) for $\text{H}_2(g)$ is 0
 ΔH_f° (kJ/mol) for $\text{C}(s)$ is 0
 ΔH_f° (kJ/mol) for $\text{O}_2(g)$ is 0



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表1

生成熱 ΔH_f° (kJ/mol) for $\text{NaHCO}_3(s)$ is -933.0

227.0 kcal/mol

$\beta = 93^\circ 19'$

0.3 1 10 20 50 70 100

1 5 10

0.3 1 10 20 50 70 100

47.7 46.9 44.0

0.3 1 10 20 50 70 100

94.4 92.5 92.5 90.2 85.5 80.0 70.0 60.0

0.3 1 10 20 50 70 100

0.06 0.20 0.30

終期 $J_s = 2.99 \text{ kcal/lmol}$

$\text{Na}_2\text{CO}_3 \cdot 4\text{NaHCO}_3$

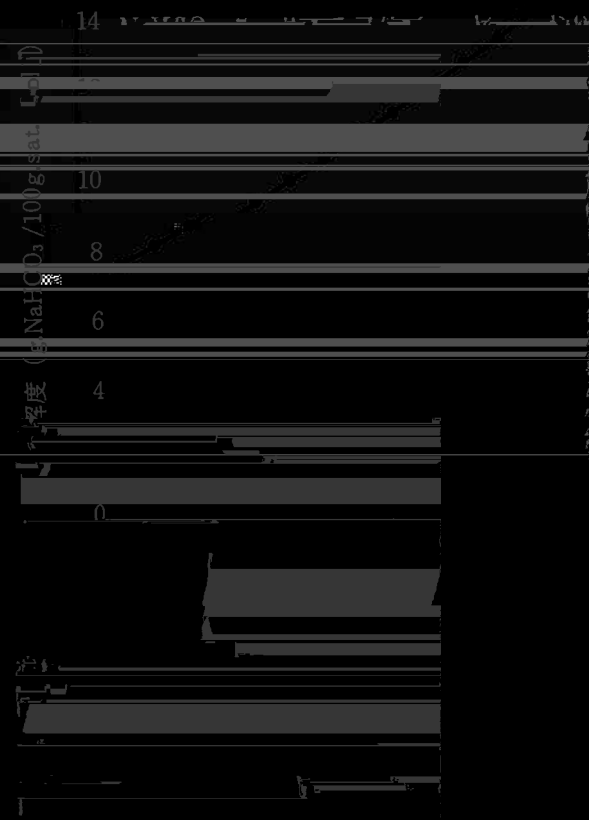
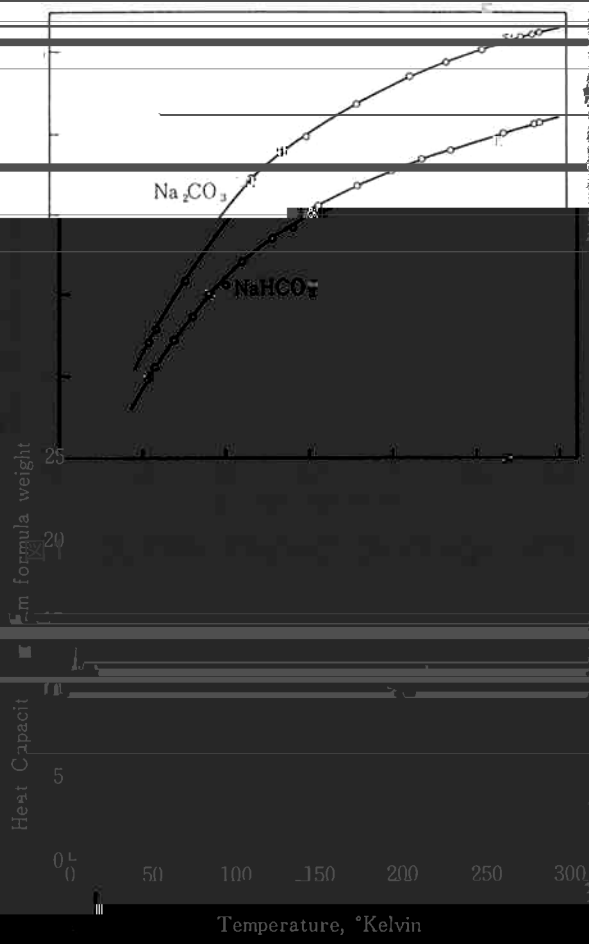
0.3 1 10 20 50 70 100

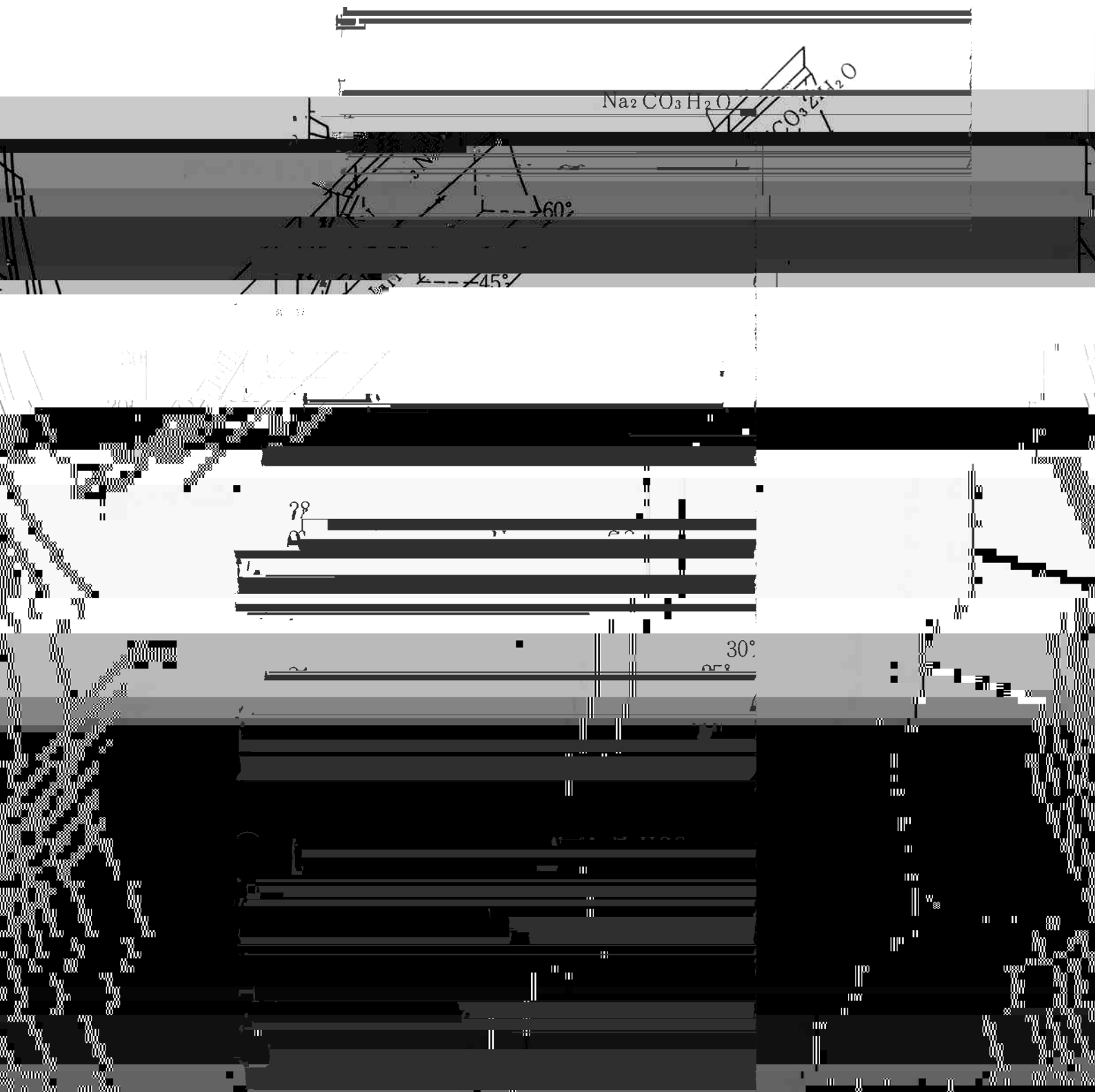
0.3 1 10 20 50 70 100

0.3 1 10 20 50 70 100

0.3 1 10 20 50 70 100

0.3 1 10 20 50 70 100

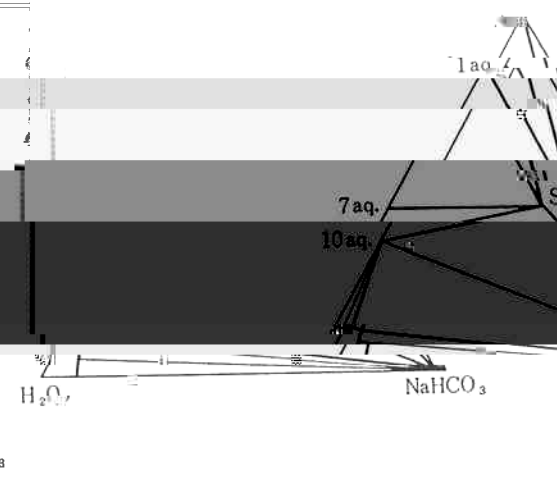
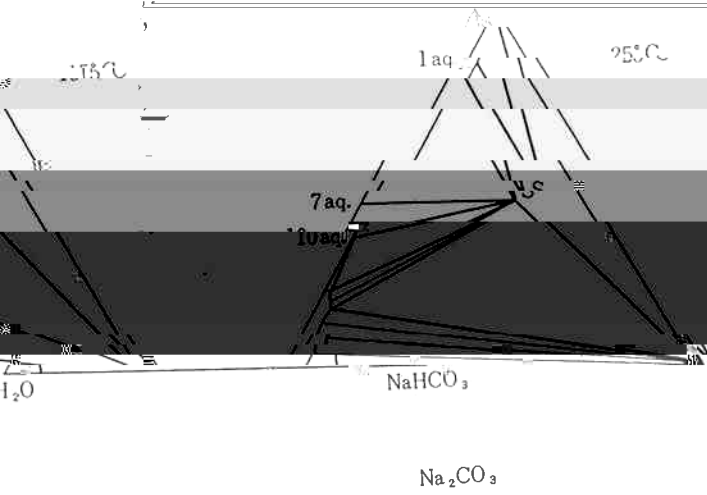




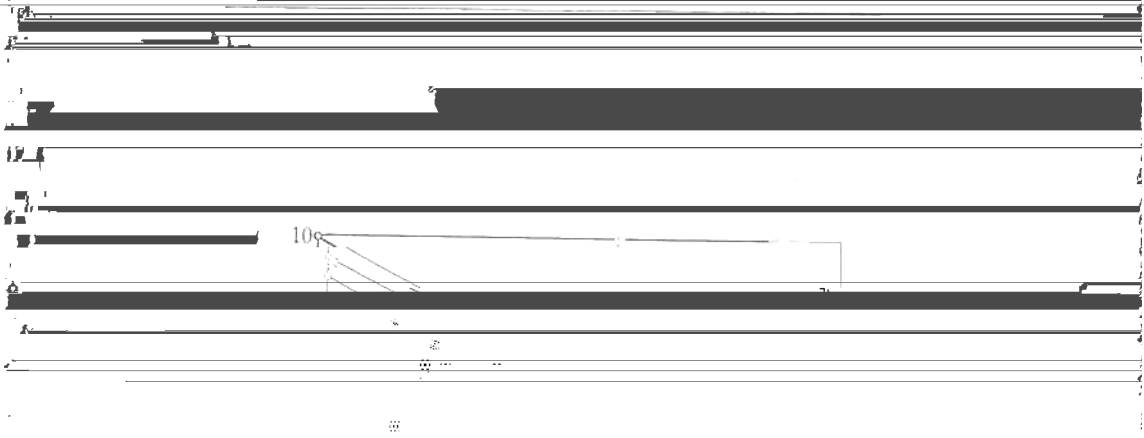
濃度 (g/100)

16
14

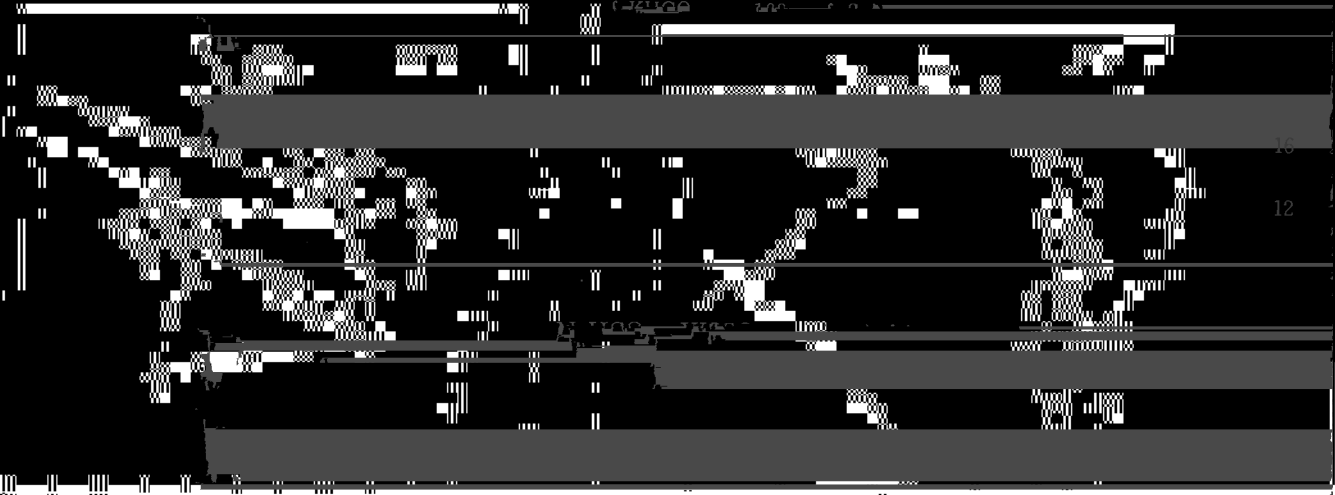
170



$\text{Na}_2\text{CO}_3 \cdot \text{H}_2\text{O}$



g. of NaHC₂O₄ per 100g. of solution



10
12

200

Moles K₂CO₃ →

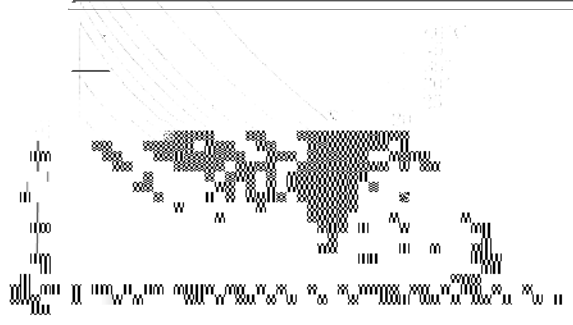


圖 7

